

STUDIES OF THE SUBSTITUTION EFFECTS ON THE ELECTRONIC PROPERTIES FOR BIPHENYL AND DERIVATIVE MOLECULES BY USING DFT METHOD

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Abstract. DFT method has been carried out to study the substitution effects of NO_2 group on the electronic properties (ionization potential, electron affinity, electronegativity, hardness, softness and electrophilicity index) and IR spectral properties of biphenyl and derivative molecules by using the B3LYP functional and the 3-21G basis set, as well as the optimization structure. The calculated values of HOMO and LUMO energies, as well as predicted by ChemBioDraw program ^1H and ^{13}C NMR spectra for the studied compounds are in a good agreement with experimental data. All properties were calculated by using Gaussian 09 program and GaussianView 5.08 program, except NMR characteristics.

Keywords: DFT calculation, electronic properties, IR and NMR spectra, biphenyl molecule.

Received: September 2015/ Revised final: January 2016/Accepted: January 2016

Introduction

Biphenyl is an aromatic hydrocarbon with a molecular formula ($\text{C}_{12}\text{H}_{10}$), which is an organic compound that forms colorless crystals. It consists of two connected phenyl rings and has D_2 symmetry [1]. Biphenyl has density 1.04 g/cm^3 , melting point 69.71°C and boiling point 255°C [2], being used in organic syntheses, food preservatives, heat transfer fluids and as an intermediate for polychlorinated biphenyls [3]. The biphenyl derivatives, such as polychlorinated biphenyls (PCB), were widely used as dielectric and coolant fluids in electrical apparatus, cutting fluids for machining operations and carbonless copy paper. 4-Aminobiphenyl has been used formerly as a rubber antioxidant and it is still used in the detection of sulfates and as a model carcinogen in mutagenicity studies and cancer research [4,5]. Many studies have been carried out on the theoretical calculations of electronic structure, ionization energy and IR spectra for azide group in 3'-azido-3'-deoxythymidine (AZT), compared to small azide compounds by the using DFT method and B3LYP/6-311G** basis set [6].

Geometry optimization has been performed and the electronic structure, total energy, electronic potential surface, UV and IR spectra were computed using TD-DFT//B3LYP method with 6-311G (d, p) basis sets for isoniazid and its derivative N-cyclopentylideneypyridine-4-carbohydrazide molecules [7]. The correlation between the crystalline structure and vibrational modes of the latter has been discussed [7]. Theoretical study of calculating the molecular structure and electronic properties of diphenylanthrazolines molecule was performed using density functional theory with B3LYP/6-31G* level [8]. Finally, a study has been carried out to highlight the substitution effects of H , NH_2 , NMe_2 , OCH_3 , CH_3 , Cl , Br , CN , NO_2 on the electronic structure and properties, such as HOMO, LUMO, energy gap and dipole moment of 2,4-domino-5-p-substituted-phenyl-6-ethyl-pyrimidines by using the DFT method and B3LYP with 6-311G (d) basis set [9].

In the light of the above-mentioned, the aim of this work was to study the influence of hydrogen atom substitution by NO_2 groups in different positions on the electronic and spectral properties for biphenyl and derivative molecules using the B3LYP functional and the 3-21G basis set.

Calculations

The DFT methods have been carried out using the Gaussian 09 program [10]. The geometric optimization of biphenyl **1** and derivative molecules **2-6** with the substitution of hydrogen atom by NO_2 groups in the different positions have been performed with a B3LYP/3-21G level in the gas phase. The hybrid functional B3LYP basis set has been shown highly successful for calculating the electronic properties, such as electronic states, ionization potentials and energy gaps, since calculation of the exchange-correlation energy and interaction effects between electrons also show the most accurate results. Figure 1 shows the geometry optimization of biphenyl **1** and derivative molecules **2-6** in the gas phase using DFT methods with B3LYP/3-21G level. The EHOMO represents the energy of highest occupied molecular orbital, while ELUMO represents the lowest unoccupied molecular orbital.

The ionization potential (IP) and electron affinity (EA) in the framework of Koopmans' theorem could be calculated from the energies of $HOMO$ and $LUMO$ as follows by Eqs.(1) and (2) [11]:

$$IP = -EHOMO \quad (1)$$

$$EA = -ELUMO \quad (2)$$

The electronegativity (χ) has been calculated by Eq.(3) [12]:

$$\chi = -\frac{IP + EA}{2} = -(EHOMO + ELUMO)/2 \quad (3)$$

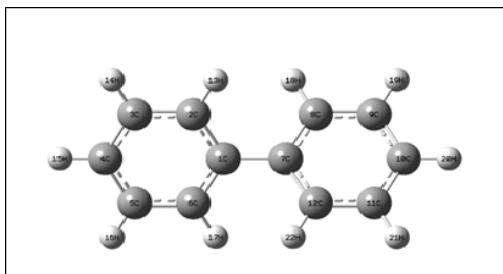
Where the hardness (η) is defined by Eq.(4) [11]:

$$\eta = \frac{IP - EA}{2} \quad (4)$$

The softness (S) and electrophilicity index (W) have been calculated by using the Eqs.(5) and (6) [13, 14]:

$$S = \frac{1}{2\eta} \quad (5)$$

$$W = \frac{\chi^2}{2\eta} \quad (6)$$



range of 1.079 - 1.085 Å and C-C bond length is at 1.373 Å and 1.498 Å, whereas C–N bond length varies from 1.338 Å to 1.476 Å. The N–O bond length of dinitrobiphenyl molecules varies in the range of 1.282 - 1.285 Å. All angles of the dinitrobiphenyl ring (C–C–C angles) vary from 117.602 to 120.220 Å^(°).

Electronic properties

All properties are computed by the DFT/B3LYP 3-21G method for biphenyl **1** and derivative molecules **2-6**. Table 2 shows the calculated energy values of *HOMO* and *LUMO* and electronic properties (values of *IP* and *EA* calculated by Eqs.(1) and (2), χ values (Eq.(3)), η values (Eq.(4)), *S* values (Eq.(5)), and *W* values calculated by Eq.(6)). It is clear from Table 2 and Figures 2-7 that the substitution of hydrogen atom by NO₂ group leads to decreasing the energy values of *HOMO* and increasing the energy values of *LUMO*. Figures 2-7 show the electronic density distribution along the rings C–C. In the different parts of dinitrobiphenyl molecules the electronic density distribution is homogeneous. It was noticed for all molecules that *HOMO* shows a bonding character and *LUMO* shows an antibonding character. While the values of *IP* are increasing, the values of *EA* are decreasing with the substitutions. The values of χ , η , *S* and *W* are increasing. The approximate values of *HOMO* and *LUMO* energies are in a good agreement with experimental values [15].

Table 1

Optimized geometrical parameters of biphenyl **1 and derivative molecules **2-6**,
R (Å) bond length, bond angles *A* (°).**

Molecules	Bond length	Our data (Å)	Bond angles	Our data (°)
1	R(C1 - C2)	1.4051	A(C2-C3-C4)	120.2203
	R(C1-C7)	1.4886	A(C2-C1-C6)	118.6205
	R(C2 – C3)	1.3949	A(C8-C7-C12)	118.6205
	R(C3 – C4)	1.3970	A(C1-C6-H17)	119.2746
	R(C2 – H13)	1.0840	A(C9-C10-C11)	120.2203
2	R(C1- C2)	1.3738	A(C2-C1-C6)	119.5174
	R(C1 – C7)	1.4793	A(C4-C3-N21)	94.8932
	R(C3- N21)	1.3386	A(C1-C6-H16)	116.6752
	R(C2 – H13)	1.0825	A(O25-N24-O26)	125.7101
	R(N24-O26)	1.2824	A(O22-N21-O23)	125.8611
3	R(C1- C2)	1.4012	A(C2-C1-C6)	117.6027
	R(C1 – C7)	1.4984	A(C2-C3-N21)	116.8991
	R(C2 – N21)	1.4613	A(C1-C6-H16)	118.5167
	R(C3 – H13)	1.0794	A(O25-N24-O26)	125.0666
	R(N24-O26)	1.2850	A(C12-N24-O25)	117.3061
4	R(C1 - C2)	1.4027	A(C2-C1-C6)	118.3123
	R(C11 – C10)	1.3970	A(C2-C3-N21)	118.9294
	R(C9- H17)	1.0835	A(C1-C6-H15)	121.7931
	R(C3-N21)	1.4675	A(O25-N24-O26)	126.1181
	R(N24-O26)	1.2809	A(C3-N21-O23)	116.9433
5	R(C1-C2)	1.4030	A(C2-C1-C6)	117.4296
	R(C11-C10)	1.3886	A(C3-C2-N21)	116.3391
	R(C9-H18)	1.0797	A(C7-C12-H20)	119.5894
	R(C10 -N21)	1.4625	A(N22-O21-N23)	125.6262
	R(N21-O22)	1.2841	A(C10-N21-O23)	117.1702
6	R(C1- C2)	1.4069	A(C2-C1-C6)	118.6886
	R(C10 – C11)	1.3928	A(C10-C9-N21)	119.0676
	R(C10 - N21)	1.4667	A(C1-C2-H13)	119.3218
	R(C2 – H13)	1.0834	A(O22-N21-O23)	125.7443
	R(N21-O22)	1.2824	A(C5-N24-O26)	117.1239

Table 2

***HOMO*, *LUMO* energies and electronic properties (*IP*, *EA*, χ , η , *S* and *W*) in (eV) units
for biphenyl **1** and derivative molecules **2-6** using DFT with B3LYP/3-21G basis set.**

Molecules	<i>HOMO</i>	<i>LUMO</i>	<i>IP</i>	<i>EA</i>	χ	η	<i>S</i>	<i>W</i>
1	-6.21309	-0.59891	6.213088	0.598914	-3.40600	2.807087	1.403543	16.28228
2	-6.02588	-2.88681	6.025876	2.886815	-4.45635	1.569530	0.784765	15.58466
3	-7.14887	-2.78532	7.148874	2.785318	-4.96710	2.181778	1.090889	26.91446
4	-7.19377	-3.23566	7.193772	3.235660	-5.21472	1.979056	0.989528	26.90850
5	-7.27432	-2.97416	7.274317	2.974162	-5.12424	2.150077	1.075039	28.22818
6	-7.34262	-2.90450	7.342616	2.904502	-5.12356	2.219057	1.109529	29.12608

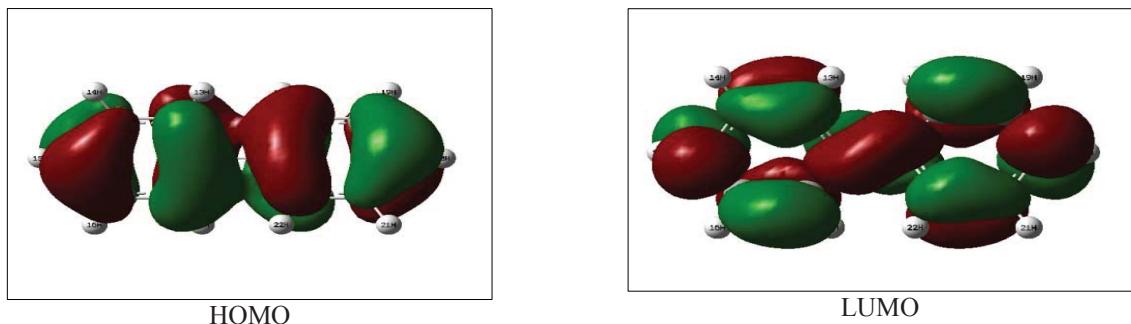


Figure 2. *HOMO* and *LUMO* shapes of the biphenyl 1 molecule using the B3LYP/3-21G method.

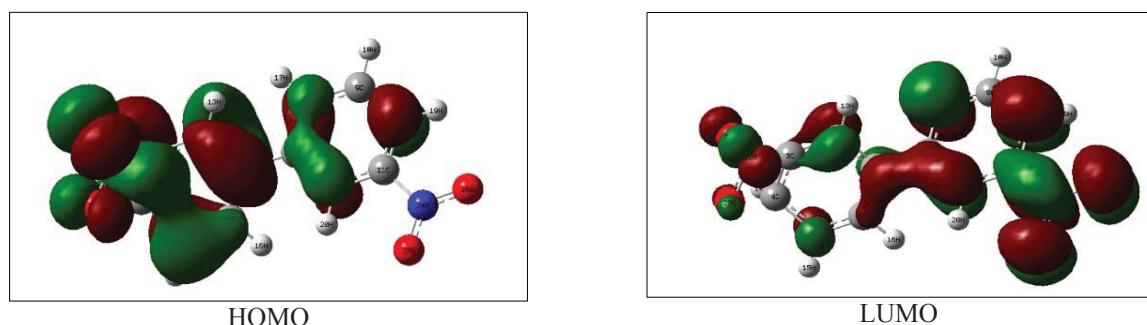


Figure 3. *HOMO* and *LUMO* shapes of the 3,11-dinitrobiphenyl 2 molecule using the B3LYP/3-21G method.

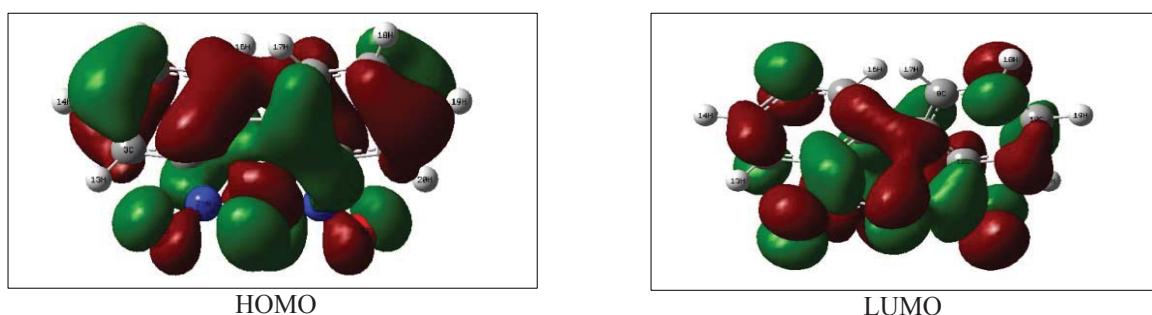


Figure 4. *HOMO* and *LUMO* shapes of 2,12-dinitrobiphenyl 3 molecule using the B3LYP/3-21G method.

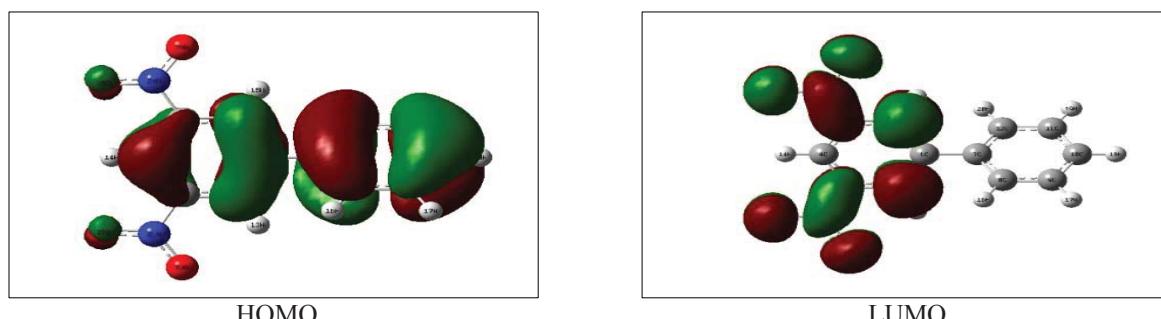


Figure 5. *HOMO* and *LUMO* shapes of 3,5-dinitrobiphenyl 4 molecule using the B3LYP/3-21G method.

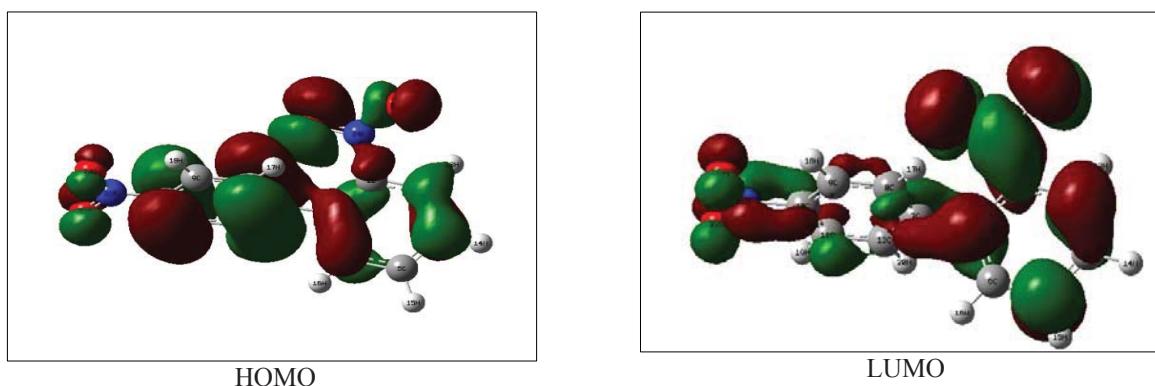


Figure 6. *HOMO* and *LUMO* shapes of 2,10-dinitrobiphenyl 5 using the B3LYP/3-21G method.

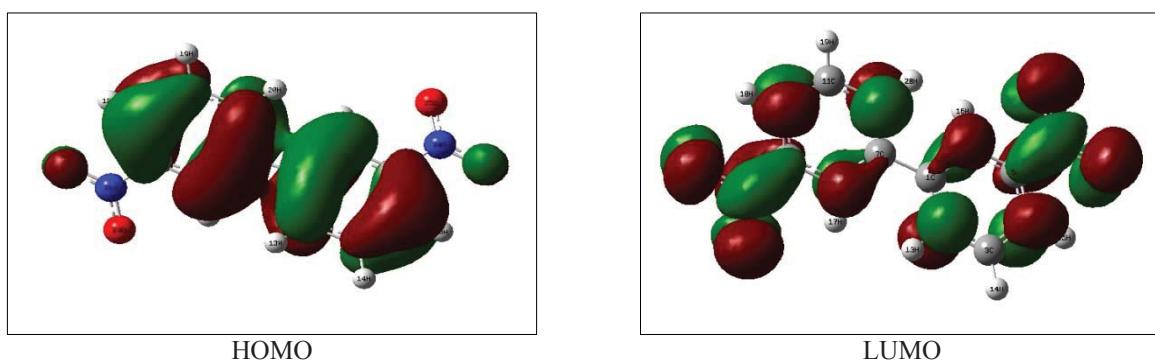


Figure 7. *HOMO* and *LUMO* shapes of 9,5-dinitrobiphenyl 6 molecule using the B3LYP/3-21G method.

IR spectra

Figures 8-10 show the computed IR spectra of biphenyl molecules **1-6**. The harmonic vibrational frequencies were calculated for the studied molecules by using B3LYP level with a 3-21G basis set. The substitution of hydrogen atoms by NO_2 leads to increasing in the number of vibrational modes. According to the rule of (3N-6), the biphenyl molecule has 60 degrees of freedom, while the dinitrobiphenyl molecules have 72 degrees of freedom. All degrees of freedom for biphenyl **1** and its derivatives **2-6** are represented by stretching, bending, rocking and scissoring mode. For a biphenyl molecule, the (C-H) symmetric stretching has been studied around 3221.33 cm^{-1} and asymmetric- at 3205.55 cm^{-1} . The in-plane vibrations have been calculated at 957.55 cm^{-1} . The out-of-plane (C-H) vibrations found at 557.30 cm^{-1} are in a good agreement with experimental values at 552.90 cm^{-1} [1]. The (C-C) stretching vibrations have been studied in the region 1634.33 cm^{-1} and (C-C-C)- at 516.16 cm^{-1} . According to Figures 8-10, in the IR spectra of dinitrobiphenyl molecules **2-6** the new peaks in the vibration modes are registered, in comparison with biphenyl molecule **1**, due to existing of stretched bond, such as (C-N), (C-H) and (N-O) for **2-6** molecules. The symmetric stretching of (C-H) has been observed in the range of 3265.99 - 3225.80 cm^{-1} and asymmetric stretching- in the region 3194.29 - 3208.81 cm^{-1} . The (C-N) stretching has been observed in the region 1600.77 - 1190.03 cm^{-1} , whilst the stretching of (C-C) bond is shown in the range of 1648.56 - 1632.25 cm^{-1} and the stretching of (N-O) bond is shown in the region 1362.28 - 1458.81 cm^{-1} . The NO_2 groups have the rocking mode appeared in the region 74.03 - 270.29 cm^{-1} , while the wagging modes appeared in the region 262.77 - 411.34 cm^{-1} and the scissoring mode appeared in the region 701.63 – 874.59 cm^{-1} .

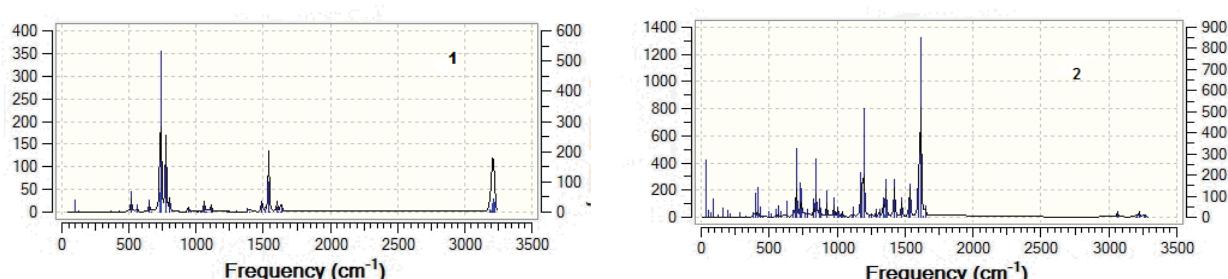


Figure 8. Calculated IR spectra of biphenyl 1 and its derivative 2.

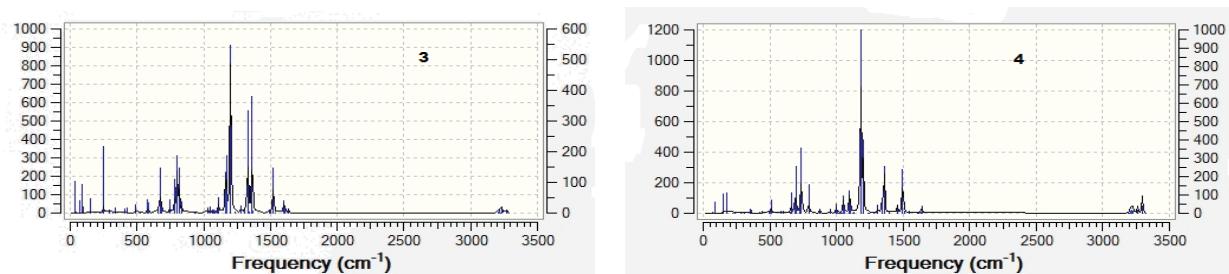


Figure 9. Calculated IR spectra of biphenyl derivatives 3 and 4.

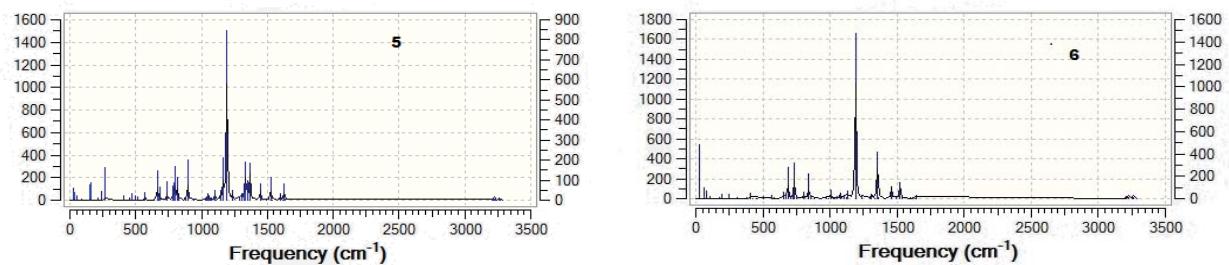
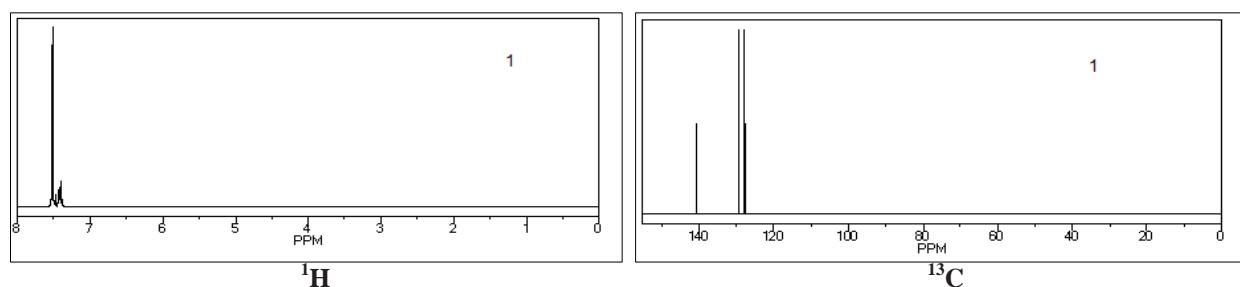
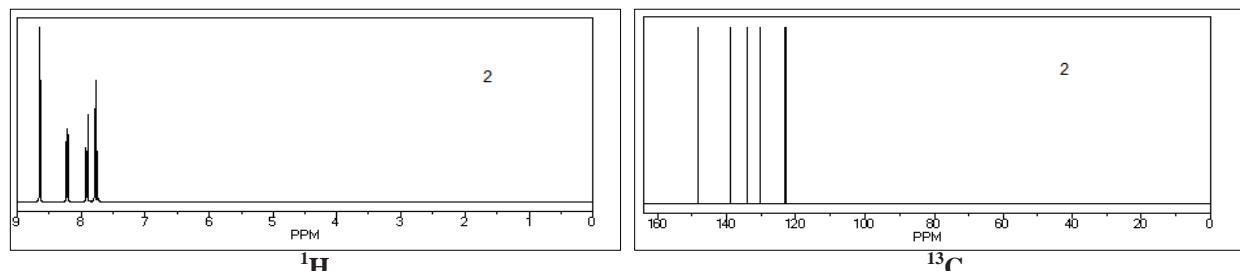


Figure 10. Calculated IR spectra of biphenyl derivatives 5 and 6.

NMR spectra

NMR spectra of ^1H and ^{13}C of biphenyl **1** and derivative molecules **2-6** (Figures 11-16) were computed by using the ChemBioDraw Program. Shielding has been considered to be one of the most effective factors that influence the chemical shift. In sites around the nucleus, where the electronic density is higher, the diamagnetic shielding becomes higher. It is clear from Figures 11-16 that the substitution of hydrogen atoms by NO_2 groups in dinitrobiphenyl molecules **2-6** leads to increasing of the value of chemical shift in ^1H NMR spectra and weak increase in chemical shift of ^{13}C nuclei. The molecules **3** and **5** have a complex ^1H NMR spectrum, because of the mutual influences between the nuclei inside the molecules. The molecule **4** has high values of chemical shift for ^1H nuclei that are found in the interval 7.4 - 9.4 ppm and ^{13}C nuclei in the interval 117.0 - 149.3 ppm. This is due to the effect of electronic density cyclic currents and electrons delocalization causing the strong shielding. The values of chemical shifts for ^1H and ^{13}C nuclei in NMR spectra for biphenyl **1** are in a good agreement with experimental values [16].

Figure 11. Predicted ^1H and ^{13}C NMR spectra for biphenyl 1.Figure 12. Predicted ^1H and ^{13}C NMR spectra for biphenyl derivative 2.

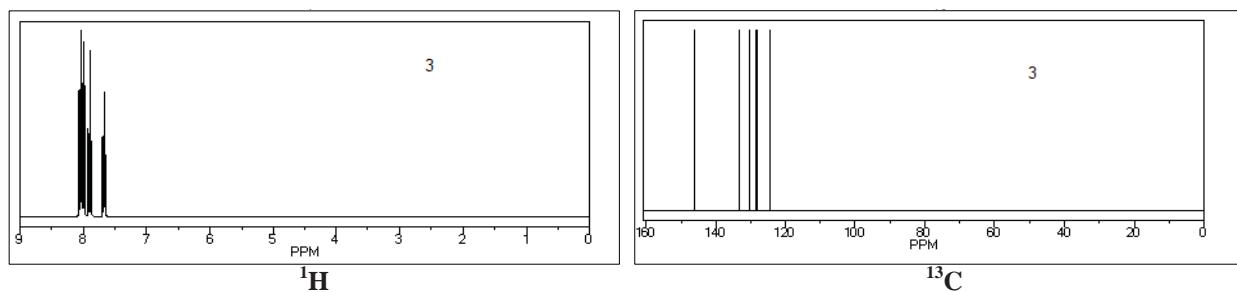


Figure 13. Predicted ¹H and ¹³C NMR spectra for biphenyl derivative 3.

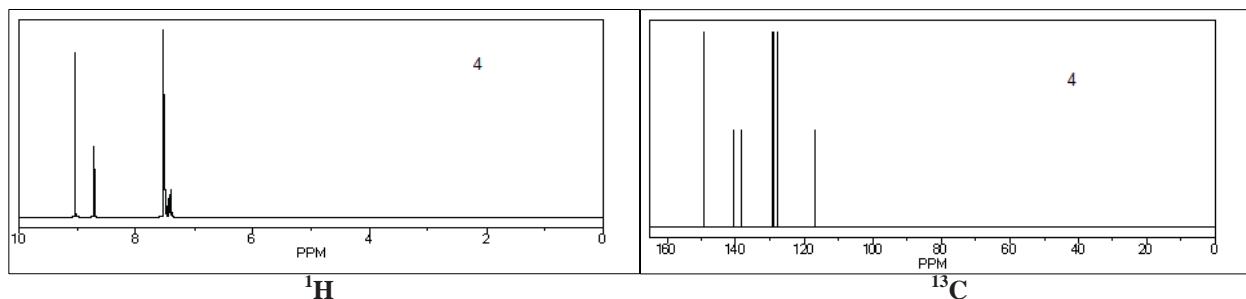


Figure 14. Predicted ¹H and ¹³C NMR spectra for biphenyl derivative 4.

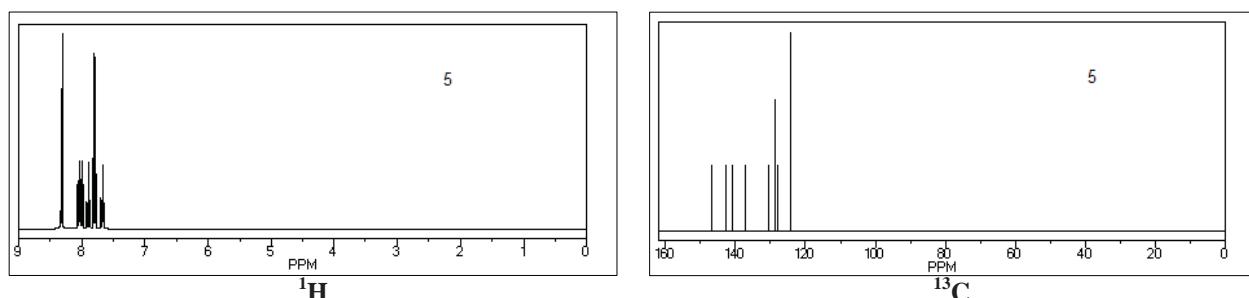


Figure 15. Predicted ¹H and ¹³C NMR spectra for biphenyl derivative 5.

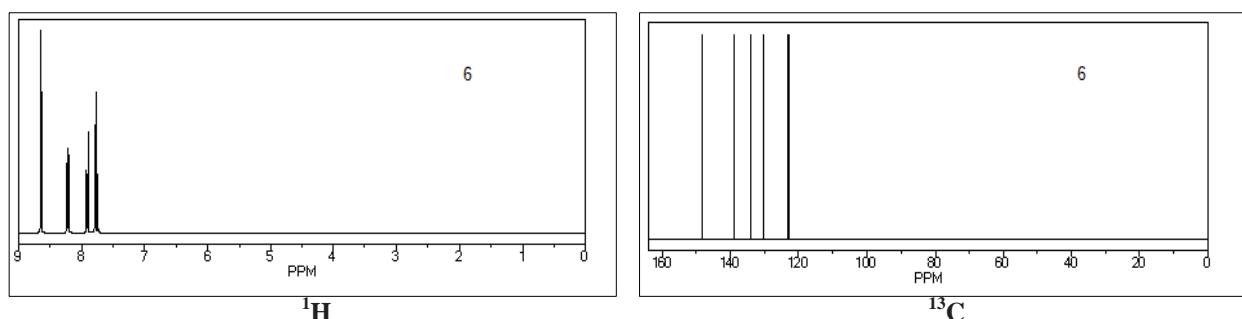


Figure 16. Predicted ¹H and ¹³C NMR spectra for biphenyl derivative 6.

Conclusions

The values of geometric optimization (bond length and its angles) for dinitrobiphenyl molecules **2-6** are slightly different in the case of substitution of hydrogen by NO₂ groups, also depending on the substituted atom's position.

The *HOMO* energy values are decreasing, while the *LUMO* energy values are increasing at the substitution of hydrogen atoms by NO₂ groups, the *HOMO* and *LUMO* energy values for biphenyl molecule **1** being in a good agreement

with experimental values. The electronic density is distributed to the different parts of dinitrobiphenyl molecules **2-6** and along the C-C rings.

The values of *IP* are increasing, while the values of the *EA* are decreasing in case of the substitution of the hydrogen atoms by NO_2 groups, while the calculated values of χ , S , η and W are increasing.

According to the IR spectra, the substitution of hydrogen atoms by NO_2 groups leads to increasing the number of vibrational modes, also the stretching vibrations of dinitrobiphenyl molecules **2-6** are increasing compeer with biphenyl molecule **1**.

According to the data of ^1H and ^{13}C NMR spectra, it was found that the substitution of hydrogen atoms by NO_2 groups causes appearance of new peaks for dinitrobiphenyl molecules **2-6**. The molecules **3** and **5** have a complex ^1H NMR spectrum, because of the mutual influences between the nuclei inside the molecules, which lead to increase in intensity of some lines with respect to the other spectrum lines. The molecule **4** has high values in chemical shift of ^1H nuclei and ^{13}C nuclei in comparison with the other dinitrobiphenyl molecules.

Acknowledgments

I gratefully thank to my supervisor Prof. Dr. Dhia Hamdi Al-Amiedy, College of Science for Women, University of Baghdad-Iraq, for his invaluable assistance during preparation of this paper.

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